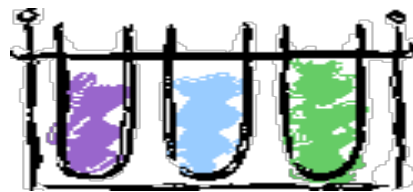


Acids and bases



By the end of this topic I will be able to:

- Recall and write the chemical formulae for the following: hydrochloric acid, sulphuric acid and sodium hydroxide
- Recall the names of common indicators and the colours they turn in acidic, basic and neutral solutions.
- Explain how to make an indicator from common materials such as flower petals.
- Classify a substance as acidic, basic or neutral by measuring its pH or from information supplied about its pH.
- Recall that pH scale ranges from 1-14 and understand the significance of pH values.
- Recall that an acid will react with a base to form a neutral salt and water.
- Explain how to extract a dissolved salt from a solution.
- Recall some everyday examples of neutralization reactions.
- Recall that an acid will react with a metal to form a metal salt and hydrogen gas.
- Explain how to test for hydrogen gas.
- Recall that carbon dioxide, a metal salt and water are produced when a metal carbonate reacts with an acid.
- Explain how to test for carbon dioxide.
- Write a balanced chemical equation for a given word equation (extension).



Nitric acid dissolving copper